

Subject-Chemistry

Activity3- Chemical Properties of Metals and Non-metals

Chapter3-Metals and Non-Metals

Class-X

Duration-40mins

Nature of the Activity- Individual

Learning Outcome-

Students will be able explore and infer the basic and acidic nature of Metallic and Non-metallic oxides.

What we have to do?

To demonstrate that Metallic oxides are basic in nature whereas Non-metallic oxides are acidic in nature

What do we require?

Magnesium ribbon, Sulphur powder, tongs, test-tube, test-tube holder, Burner, Blue litmus solution and Red litmus solution

What do we need to know?

Metals react with oxygen to form metallic oxides whereas Non-metals react with oxygen to form non-metallic oxides

How will you proceed?

➤ Teacher's Activity

Teacher will brief the activity and make the necessary arrangements for the activity.

➤ Student's activity



Take a Magnesium ribbon and some Sulphur powder.

Hold the magnesium ribbon with tongs and burn them on the burner.

Collect the ashes so formed and dissolve them in water.

Test the resultant solution with both Red Litmus Solution and Blue Litmus Solution

Record your observations/.

Heat a small amount of Sulphur powder in a clean and dry test tubes and invert another test tube over it.

What do we observe?

Sample	Red Litmus Solution	Blue Litmus Solution	Nature of the oxide
Metallic Oxide			
Non-Metallic Oxide			

What you have learnt?

Metallic Oxides are _____ in nature

Non-metallic Oxides are _____ in nature

What would you be interested to know?

1. What are amphoteric oxides?
2. How does a Metallic oxide react with an acid?
3. How does a Non-metallic oxide react with a base?

