

ACIDS, BASES AND SALTS



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FAMILY OF SALTS



WRITE THE FORMULAE OF THE SALTS :

- POTASSIUM SULPHATE
- SODIUM SULPHATE
- CALCIUM SULPHATE
- MAGNESIUM SULPHATE
- COPPER SULPHATE
- SODIUM CHLORIDE
- SODIUM NITRATE
- SODIUM CARBONATE
- AMMONIUM CHLORIDE

FAMILY OF SALTS



- Salts having the same positive or negative radicals are said to have a common family.
- E.g. Family of sodium salts- Sodium chloride and Sodium sulphate
- E.g. Family of chloride salts - Sodium chloride and Potassium chloride

pH OF SALTS



- Strong acid + Strong base $\rightarrow$ Neutral Salt
- Strong acid + Weak base $\rightarrow$ Acidic Salt
(pH value less than 7)
- Weak acid + Strong base $\rightarrow$ Basic Salt
(pH value more than 7)

COMMON SALT



- THE SALT FORMED BY THE COMBINATION OF HYDROCHLORIC ACID AND SODIUM HYDROXIDE SOLUTION IS CALLED SODIUM CHLORIDE.
- THIS IS THE COMMON SALT THAT WE USE IN OUR HOME.
- SEA WATER CONTAINS MANY SALTS DISSOLVED IN IT. SODIUM CHLORIDE IS SEPARATED FROM THESE SALTS.
- ROCK SALT – THESE ARE THE LARGE DEPOSITS OF SOLID SALT THAT ARE FOUND IN SEVERAL PARTS OF THE WORLD. THEY ARE OFTEN BROWN DUE TO IMPURITIES.



CHLOR-ALKALI PROCESS



- The aqueous solution of sodium chloride is called Brine.
- When Electricity is passed through aqueous solution of sodium chloride , it decomposes to form sodium hydroxide. Chlorine gas is given off at anode and hydrogen gas at the cathode.
- All the three products (SODIUM HYDROXIDE , CHLORINE GAS , HYDROGEN GAS) have different uses.

