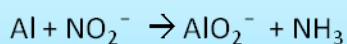
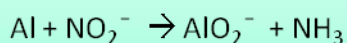


REDOX REACTIONS

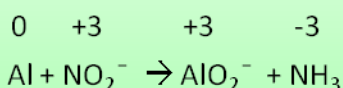
Balancing Redox Reaction in Basic Medium



Step 1: Balance all atoms except H and O (In this case all are balanced)



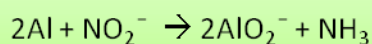
Step 2: Write the Oxidation Numbers (ON) to identify oxidation and reduction



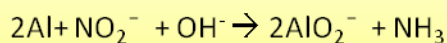
(Oxidation Number of Al increases by 3 while oxidation number of N decreases by 6)



Step 3: Equate the increase and decrease in ON (In this case N undergoes decrease in ON by 6 so make Al atoms 2 on both sides so that total increase in ON is also by 6)



Step 4: Balance charges (adding OH^- as this is basic medium)



Step 6: Add water to balance Oxygen (Hydrogen gets balanced automatically)

